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The oxidation number of a free element is always 0. The atoms in He and N 2, for example, have oxidation numbers of 0. The oxidation number of a monatomic ion equals the charge of the ion. For example, the oxidation number of Na + is +1; the oxidation number of N 3-is -3. The usual oxidation number of hydrogen is +1.

Rules for Assigning Oxidation Numbers - ThoughtCo

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There are sets of rules for assigning oxidation numbers. The total sum of the oxidation numbers has to equal the total charge of the molecule or ion. O is almost always -2, unless it is in peroxide.

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Assigning Oxidation Numbers. The oxidation number is a positive or negative number that is assigned to an atom to indicate its degree of oxidation or reduction. In oxidation-reduction processes, the driving force for chemical change is in the exchange of electrons between chemical species. A series of rules have been developed to help us.

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